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*A Bibliography of Sources of
Thermodynamic Data for the Systems:
 $CO_2 + NH_3 + H_2O$, $CO_2 + H_2S + H_2O$,
 $H_2S + NH_3 + H_2O$, and $CO_2 + NH_3 + H_2S + H_2O$*

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A Bibliography of Sources of Thermodynamic Data for the Systems: $\text{CO}_2 + \text{NH}_3 + \text{H}_2\text{O}$, $\text{CO}_2 + \text{H}_2\text{S} + \text{H}_2\text{O}$, $\text{H}_2\text{S} + \text{NH}_3 + \text{H}_2\text{O}$, and $\text{CO}_2 + \text{NH}_3 + \text{H}_2\text{S} + \text{H}_2\text{O}$

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ABSTRACT

Contained herein is a bibliography of sources of experimental and correlated thermodynamic data for the systems $\text{CO}_2 + \text{NH}_3 + \text{H}_2\text{O}$, $\text{CO}_2 + \text{H}_2\text{S} + \text{H}_2\text{O}$, $\text{H}_2\text{S} + \text{NH}_3 + \text{H}_2\text{O}$, and $\text{CO}_2 + \text{NH}_3 + \text{H}_2\text{S} + \text{H}_2\text{O}$. The types of data in this bibliography include all types of equilibrium data, including both equilibria in solution and vapor-liquid equilibrium data, enthalpies, heat capacities, and densities. There are 215 references cited.

KEY WORDS: ammonia; ammonium carbamate; bibliography; carbon dioxide; enthalpy; equilibrium constants; heat capacity; hydrogen sulfide; sour water; thermodynamics; urea; vapor-liquid equilibria; water.

INTRODUCTION

The thermodynamics of mixtures of $\text{CO}_2 + \text{NH}_3 + \text{H}_2\text{O}$, $\text{CO}_2 + \text{H}_2\text{S} + \text{H}_2\text{O}$, $\text{NH}_3 + \text{H}_2\text{S} + \text{H}_2\text{O}$, and $\text{CO}_2 + \text{NH}_3 + \text{H}_2\text{S} + \text{H}_2\text{O}$ are of importance for a variety of industrial applications which include the synthesis of urea, the utilization of sour water systems, gas production, and environmental concerns. The purpose of this bibliography is to identify papers which contain either experimental thermodynamic data for these systems or reviews or correlations of this data. The thermodynamic properties of interest include all types of equilibrium data, including both equilibria in solution and vapor-liquid equilibrium data, enthalpies, heat capacities, and densities. The principal species in aqueous solutions are CO_2^0 , CO_3^{2-} , HCO_3^- , H_2CO_3^0 , NH_3^0 , NH_4^+ , HS^- , S^{2-} , $\text{CO}(\text{NH}_2)_2$, and ammonium carbamate, $\text{NH}_2\text{COONH}_4$. Our search of the literature was based upon a search of the following sources: a computer aided one of Chemical Abstracts from 1967 to 1983, the Bulletin of Chemical Thermodynamics [1] from 1960 to 1981, the files of the Chemical Thermodynamics Data Center at the National Bureau of Standards, and, finally, of the references found in the papers identified in the search itself. The authors would appreciate learning of any papers which we have missed in our search.

The papers are listed alphabetically by first author. Each bibliographic citation includes a brief reference citation (the year and three letters from the names of the first two authors), the authors names, the title of the article and the source. In one instance only an abstract was available and there we have also given the Chemical Abstracts citation. The last column contains a capital letter(s) which serves to identify which systems are found in the paper. The letter codes used are: (A) $\text{CO}_2 + \text{H}_2\text{S} + \text{H}_2\text{O}$, (B) $\text{CO}_2 + \text{NH}_3 + \text{H}_2\text{O}$, (C) $\text{H}_2\text{S} + \text{NH}_3 + \text{H}_2\text{O}$, and (D) $\text{CO}_2 + \text{H}_2\text{S} + \text{NH}_3 + \text{H}_2\text{O}$.

Each paper has been annotated to show the type of data, the temperature range and, if appropriate, the pressure, solution composition, and pH. The property codes which we have used are from the Bulletin of Chemical Thermodynamics [1]; they are reproduced in Table I. In specifying the compositions of ternary systems, we have frequently used the quantities L and W, which are defined to be the mole ratios NH_3/CO_2 and $\text{H}_2\text{O}/\text{CO}_2$, respectively. Note that, using existing conventions, the value of W can be negative [2]. We have attempted to adhere to the journal abbreviations used in the Chemical Abstracts Service Source Index [3].

The Japanese, Russian, and Polish titles have been translated into English. These translations come, by preference, from the papers themselves or from Chemical Abstracts. The titles from the latter source may be condensations of the originals.

This bibliography was sponsored by the Design Institute for Physical Property Data (DIPPR) of the American Institute of Chemical Engineers. Bibliographies on the binary systems $\text{CO}_2 + \text{H}_2\text{O}$, $\text{NH}_3 + \text{H}_2\text{O}$, and $\text{H}_2\text{S} + \text{H}_2\text{O}$ have also been compiled for DIPPR under the direction of Drs. David Garvin, David Smith-Magowan, and Bert R. Staples. These bibliographies have been used in developing the present listing, but those papers which treat only the binaries have not been repeated here. We thank Mr. T. B. Selover of the Standard Oil Company for bringing several references to our attention for inclusion in this bibliography, Dr. Hideo Okabe for his assistance with the papers written in Japanese, and Dr. Ewa Gajewski for her help with the papers in Polish.

REFERENCES

- [1] Bulletin of Chemical Thermodynamic, R. D. Freeman, editor,
Thermochemistry, Inc., Stillwater, OK, 1977 to 1981.
- [2] Durisch, W.; Lemkowitz, S. M.; van den Berg, P. J.; Chimia; *34, 314
(1980); "Constituent and Component Measurements and Calculations of the
Vapor/Liquid Equilibrium of the Ternary System Carbon
Dioxide-Ammonia-Water under Urea Synthesis Conditions."
- [3] Chemical Abstracts Service Source Index: 1907-1979 Cumulative, in two
volumes; American Chemical Society, Washington, DC, 1979.

TABLE I

<u>Property Symbol</u>		<u>Description of Properties</u>
<u>Category</u>	<u>Subgroup</u>	
D		<u>Derived Quantities of Special Interest (not listed elsewhere)</u>
	Da	Activity, fugacity
	Dm	Partial molar quantities
	Dx	Excess functions for mixtures
H		<u>Enthalpy Changes for Processes, by Calorimetry</u>
	Hc	for combustion in O ₂ or F ₂
	Hr	for other chemical reactions
	Hm	for mixing: solution, dilution, etc.
	Hp	for phase transitions, fusion, vaporization
	Hs	for surface processes: adsorption, desorption, etc.
K		<u>Reaction Equilibria and Related Data</u>
	Kd	Dissociation/decomposition pressures and derived enthalpy/entropy changes.
	Ke	Electrochemical cell potentials, etc. and derived enthalpy/entropy changes.
	Kk	Equilibrium constants for chemical reactions and derived enthalpy/entropy changes.
M		<u>Thermodynamic Quantities Calculated from Molecular Parameters</u>
	Md	Data (e.g., structural or spectroscopic) for molecular parameters; atomic energy levels; ionization potentials.
	Mi	Ideal gases: thermodynamic functions (e.g., tabulations).
	Mg	Real gases: intermolecular potentials, derived equations of state
	Mm	Mixtures
	Mx	Crystal, solid and liquid states.
	Mb	Bond energies; non-bonded interactions.

P		<u>Phase Equilibria</u> Temperatures: freezing point, boiling point, triple point, other phase transitions. Vapor pressure and derived quantities for vaporization and/or sublimation, pure substances Vapor/liquid equilibria and related phase diagrams, <u>mixtures</u> Condensed phase equilibria: solubility, freezing points, phase diagrams. Osmotic pressure/membrane equilibria Surface phenomena: surface tension, surface energy, adsorption, etc.
Q	Pt Pp Pv Px Po Ps Q1 Qh Qg	Thermal Properties for Non-Reacting Systems, by Calorimetry Condensed phase, $T \leq 400$ K: heat capacity, enthalpy, entropy, etc. Condensed phase, $T \geq 400$ K: heat capacity, enthalpy, entropy, etc. Gas phase: heat capacity, enthalpy, entropy etc. as $f(T,P)$.
V	Vg Vc Vx Vt	<u>Volume as $f(T,P)$; Empirical Equations of State</u> Gases: PVT, and related data Critical state properties Condensed phases: compressibility, thermal expansivity Tables and charts of data, e.g., Mollier diagrams.
X	Xd Xv Xr	<u>Physical Properties of a Single Phase</u> Density Viscosity Refractive index
Z	Za Zc Ze Zr	<u>Compilations, Correlations and Reviews</u> Analysis of experimental data and of errors Empirical Correlations Evaluations and compilations Reviews
Physical States:	(c)ystal, solid (amorp)hous (liq)uid (g)as	(aq)ueous (nonaq)ueous, includes fused salts, solid solutions (Ads)orbed

39ARD:	Ardeeva, V. A.; "Partial pressures of NH_3 , CO_2 and H_2O in liquids in reflux-condensors"; Zh. Khim. Prom.; *16, 26 (1939)	Pv, 70-90 °C.	B
47BAD/WIL:	Badger, E. H. M.; Wilson, D. S.; "Vapour pressures of ammonia and carbon dioxide in equilibrium with aqueous solutions. Part VI"; J. Soc. Chem. Ind. London; *66, 84 (1947)	Zc, Kk, Pv, 0-100 °C.	B
63BAR:	Baranski, A.; "Study of the kinetics of the synthesis and decomposition of urea. II. Kinetics of the decomposition of urea"; Chem. Stosow.; *4, 567 (1963)	Kk, 150-180 °C.	B
64BAR:	Baranski, A.; "Study of the kinetics of the synthesis and decomposition of urea. III. Kinetics of the decomposition of urea"; Chem. Stosow.; *3A, 281 (1964)	Kk, 120-150 °C.	B
78BEU/REN:	Beutier, D.; Renon, H.; "Representation of NH_3 - H_2S - H_2O , NH_3 - CO_2 - H_2O , and NH_3 - SO_2 - H_2O vapor-liquid equilibria"; Ind. Eng. Chem. Process Des. Dev.; *17, 220 (1978)	Zc, Kk, Da, Pv, 0-100 °C.	B,C
63BLA/BAR:	Blasiak, E.; Baranski, A.; Matuszewski, Z.; "On the kinetics of the reactions of synthesis and decomposition of urea"; Bull. Acad. Pol. Sci.; *11, 261 (1963)	Kk, 150-180 °C.	B
80BLA/SWA:	Blauwhoff, P. M. M.; van Swaaij, W. P. M.; Enschede, N.; "Gas-liquid equilibria between H_2S , CO_2 and aqueous amine solutions"; EFCE Pub. Ser. 80, Phase Equil. Fluid Prop. Chem. Ind.; *11, 78 (1980)	Zc, Pv, 40-66 °C.	A
37BOL/LEM:	Bolotov, V. A.; Leman, V. R.; Popova, A. N.; "Synthesis of ammonia-urea fluids in large-scale laboratory installations in continuous operation"; Khim. Promst. (Moscow); *14, 707 (1937)	Kk, Pv, 160-185 °C.	B

37BOL/LEM2:	Bolotov, V. A.; Leman, V. R.; Popova, A. N.; Shanoshchnikov, V. S.; "Synthesis of urea from ammonia and carbon dioxide at installations in continuous operation"; Khim Promst. (Moscow); *14, 1693 (1937)	Kk, Pv, 160-185 °C.	B
40BOL/LEM:	Bolotov, V. A.; Leman, V. R.; "Study of Partial Pressures of NH ₃ , CO ₂ , and H ₂ O over aqueous solutions of ammonium carbamate at 140 - 200 °C"; J. Chem. Ind. (USSR); pp. 28-32 (1940).	Kk, 140-200 °C	B
34BOL/POP:	Bolotov, B. A.; Popova, A. N.; "The synthesis of urea from ammonia and carbon dioxide"; Zh. Khim. Promsti.; *11, 32 (1934)	Kk, Pv, 155-270 °C.	B
37BOL/POP:	Bolotov, V. A.; Popova, A. N.; Sokolova, Yu K.; "Synthesis of urea from ammonia in carbonic acid in an abundance of ammonia"; Khim. Promst. (Moscow); *14, 631 (1937)	Kk, Pv, 135-175 °C.	B
37BOL/TUG:	Bolotov, V. A.; Tugai, D. G.; "Regeneration of gases after the synthesis of urea from ammonia and carbonic acid"; Zh. Khim. Prom.; *14, 991 (1937)	Kk, Pv, 100-140 °C.	B
75BRO/LEM:	Broers, J. N.; Lemkowitz, S. M.; van den Berg, P. J.; "Densities of urea-ammonia-water-carbon dioxide solutions in chemical equilibrium at and above urea synthesis conditions"; J. Appl. Chem. Biotechnol.; *25, 769 (1975)	Pv, Vc, Vq, Xd, 150-260 °C	B
10BUC:	Buch, K.; "Die Hydrolyse der Ammoniumsalze fluchtiger Sauren"; Z. Physik. Chim. (Leipzig); *70, 66 (1910)	Kk, Pv, 18-25 °C.	B
75BUC/RAT:	Buck, A.; Rathgeb, K.; "Gleichgewicht des Systems CO ₂ -NH ₃ -H ₂ O Harnstoff im Gebiet kleiner Wassergehalte"; Helv. Chim. Acta; *58, 81 (1975)	Kk, Pv, 160-180 °C.	B

65BUN/KHA:	Bunakov, N. G.; Kharlampovich, G. D.; "The solubility of acidic gases (carbon dioxide and hydrogen sulfide) in aqueous ammonium orthophosphate solutions"; J. Appl. Chem. USSR (Eng. Trans.); *38, 1879 (1965)	Kk, Pv, 293-363 K, H_3PO_4 present in solutions studied.	D
81CHE:	Chen, C. C.; "Computer simulation of chemical processes with electrolytes" Dissertation, Massachusetts Institute of Technology, Cambridge, Massachusetts (1981).	Zc, Kk, 293-413 K	B, C, D
79CHE/BRI:	Chen, C. C.; Britt, H. I.; Boston, J. F.; Evans, L. B.; "Extension and application of the Pitzer equation for vapor-liquid equilibrium of aqueous electrolyte systems with molecular solutes"; AIChE J.; *25, 820 (1979)	Zc, Kk, 293-413 K	B
80CHE/BRI:	Chen, C. C.; Britt, H. I.; Boston, J. F.; "Two new activity coefficient models for the vapor liquid equilibrium of electrolyte systems"; Evans, L. B.; Am. Chem. Soc. Symp. Ser. No. 133, pp. 63-89 (1980)	Zc, Da, Kk, Pv, 20-25 °C.	B
71CHE:	Chernen'kaya, E. I.; "Experimental determination of the specific heats of aqueous solutions of NH_4HCO_3 , $NaHCO_3$, Na_2CO_3 , NH_3 , and of liquors of the soda industry at 25 °C"; J. Appl. Chem. USSR (Eng. Trans.); *44, 1562 (1971)	Q1, 25 °C.	B
72CHE/BRA:	Chernenkaya, E. I.; Bratash, E. G.; "Experimental determination of the specific heats of aqueous solutions of Na_2CO_3 , $NaHCO_3$, NH_4HCO_3 , and of liquors of the soda industry at 35 and 50 °C"; J. Appl. Chem. USSR (Eng. Trans.); *45, 2325 (1972)	Q1, 25-50 °C.	B
75CHE/BRA:	Chernen'kaya, E. I.; Bratash, E. G.; "Calculation of specific heats of aqueous salt systems over a wide temperature range"; J. Appl. Chem. USSR (Eng. Trans.); *48, 1910 (1975)	Zc, Q1, 25-90 °C, empirical method for the calculation of Cp of aqueous salt sols.	B

81CHE/TRO:	Chertokova, L. V.; Troyanov, S. I.; "Equilibrium method for determination of dissolved gases"; Vulkanolog. Seismolog.; *6, 115 (1981)	Kk, Pv, 33-130 °C.	A
33CLA/GAD:	Clark, K. G.; Gaddy, V. L.; Rist, C. E.; "Equilibria in the ammonium carbamate urea-water system"; Ind. Eng. Chem.; *25, 1092 (1933)	Kk, 135-200 °C, equilibrium study of the conversion of ammonium carbamate into urea and water.	B
74CZA/HAB:	Czarnecki, J.; Haber, J.; Pawlikowski-Czubak, J.; Pomianowski, A.; "Thermodynamic equilibria in the system SO ₂ , NH ₃ , H ₂ O, and CO ₂ "; Z. Anorg. Allg. Chem.; *410, 213 (1974)	Kk, Pv, 25-60 °C.	B
62DAN/FRE:	D'Ans, J.; Freund, H. E.; "Anorganische Stoffe in Wasser, ternäre Systeme"; Landolt-Bornstein: Numerical Data and Functional Relationships in Science and Technology, 6th Edition, Volume II, Part 2, page 3-225, Springer Verlag, Berlin (1962)	Zc, Px, -95 to + 34 °C	B
80DAV/GOO:	Dave, S. M.; Goomer, N. C.; Sadhukan, H. K.; "Molecular distribution of deuterated hydrogen, water, hydrogen sulphide, methyl amine and ammonia"; Ind. J. Chem.; *19A, 733 (1980)	Zc, K, 193-883 K, calculated values of molecular distributions of deuterated derivatives of H ₂ , H ₂ O, H ₂ S, CH ₃ NH ₂ , and NH ₃ .	C
47DRY:	Dryden, I. G. C.; "Equilibrium between gaseous carbon dioxide and hydrogen sulphide and solutions of alkali carbonates, bicarbonates, and hydrosulphides. Part I. Potassium salts"; J. Soc. Chem. Ind. London; *66, 59 (1947)	Kk, 0-60 °C, 0.2 to 2.0 N.	A
78DUR:	Durisch, W.; "Experimentelle and thermodynamische Untersuchung des Siedegleichgewichts des Systems CO ₂ /NH ₃ /H ₂ O unter Harnstoffsynthese - Bedingungen" Dissertation, Eidgenössischen Technischen Hochschule, Zürich, Switzerland (1978)	Px, 150-220 °C, 70 to 280 bar	B

81DUR:	Durisch, W.; "Siedegleichgewicht chemisch reagierender Mehrstoffsysteme unter hohem Druck und bei hoher Temperatur"; <i>Chimia</i> ; *35, 19 (1981)	Zc, Pv, 180 °C.	B
79DUR/BER:	Durisch, W.; van den Berg, P. J.; "Die Dichte gasförmiger Kohlendioxid-Ammoniak-Wasser-Gemische unter Harnstoffsynthese-Bedingungen"; <i>Chimia</i> ; *33, 200 (1979)	Pv, Vc, Vg, Xd, 160–200 °C, CO ₂ : 40 to 65 mole percent, NH ₃ : 35 to 55 mole percent.	B
75DUR/BUC:	Durisch, V. W.; Buck, A.; "Zusammensetzung und Dichte der Harnstoffschmelze und der koexistierenden Gasphase"; <i>Chimia</i> ; *29, 436 (1975)	Kk, 170 °C, L = 2, W = 0.	B
77DUR/BUC:	Durisch, W.; Buck, A.; Luder, J.; "Der Gleichgewichtsdruck des Systems CO ₂ /NH ₃ /H ₂ O bei Harnstoffsynthese-Bedingungen"; <i>Chimia</i> ; *31, 296 (1977)	Pv, Xd, Vg, 160–200 °C, W = 0 to -0.6, L = 2.	B
79DUR/BUC:	Durisch, W.; Buck, A.; Lemowitz, S. M.; van den Berg, P. J.; "Vapour/liquid equilibria under high pressure with chemical reactions"; <i>Delft. Prog. Rep.</i> ; *4, 159 (1979)	Pv, Vg, Xd, 160–180 °C, L = 2, -0.6 < W < 0.	B
79DUR/BUC2:	Durisch, W.; Buck, A.; Lemkowitz, S. M.; van den Berg, P. J.; "Experimental investigation of the vapour-liquid equilibrium of the carbon dioxide ammonia-water system at urea synthesis conditions"; <i>Chimia</i> ; *33, 293 (1979)	Pv, Vg, Xd, 160–180 °C, L = 2, -0.6 < W < 0.	B
80DUR/LEM:	Durisch, W.; Lemkowitz, S. M.; van den Berg, P. J.; "Constituent and component measurements and calculations of the vapour/liquid equilibrium of the ternary system carbon dioxide ammonia-water under urea synthesis conditions"; <i>Chimia</i> ; *34, 314 (1980)	Zc, Kk, Pv, 160–180 °C, L = 2, -0.6 < W < 0.	B
78EDW/MAU:	Edwards, T. J.; Maurer, G.; Newman, J.; Prausnitz, J. M.; "Vapor-liquid equilibria in multi-component aqueous solutions of volatile weak electrolytes"; <i>AIChE J.</i> ; *24, 966 (1978)	Zc, Da, Pv, 0–170 °C, 0 < I < 6 mol kg ⁻¹ .	D

75EDW/NEW:	Edwards, T. J.; Newman, J.; Prausnitz, J. M.; "Thermodynamics of aqueous solutions containing volatile weak electrolytes"; AIChE J.; #21, 248 (1975)	Zc, Da, Kk, Pv, 0-100 °C, 10-4 to 2 mol Kg-1.	D
62EFR/LEO:	Efremova, G. D.; Leont'eva, G. G.; "Compressibility of mixtures of ammonia and carbon dioxide and the equilibrium reaction for the synthesis of ammonia"; Khim. Promst. (Moscow); #10, 742 (1962)	Kk, Pv, 200-400 °C.	B
54EGA/VAN:	Egalon, R.; Vanhille, R.; Willemyn, M.; "Tensions de vapeur de l'ammoniac et de l'anhydride carbonique en equilibre avec les solutions de carbonate et de bicarbonate d'ammonium"; Ind. Chim. (Paris); #41, 293 (1954)	Kk, Pv 21-52 °C.	B
79EKG/STA:	Ekgauz, V. I.; Starodubtsev, A. N.; Nazarov, V. G.; Zabello, L. A.; "Equilibrium distribution of ammonia, hydrogen sulfide, and carbon dioxide after absorption from water"; Vopr. Tekhnol. Ulavlivan. Pererab. Prod. Koksovaniya; #8, 16 (1979)	Kk, Pv, 30 °C.	C
04FED:	Fedotieff, P. P.; "Der Ammoniaksodaprozess vom Standpunkte der Phasenlehre"; Z. Phys. Chem. (Leipzig); #49, 162 (1904)	Px, 0-45 °C, in NaCl sols.	B
11FIC/BEC:	Fichter, F.; Becker, B.; "Über die Bildung von Harnstoff durch Erhitzen von Ammonium-carbaminat"; Ber. Bunsenges. Phys. Chem.; #44, 3433 (1911)	Kk, 115-150 °C.	B
48FRE:	Frejacques, M.; "Les bases theoriques de la synthese industrielle de l'uree"; Chim. Ind.; #60, 22 (1948)	Hr, Kk, Pv, -20 to 150 °C.	B
57FRO:	Frolich, G. J.; "Vapor liquid equilibria of aqueous systems containing ammonia and carbon dioxide"; Thesis, Polytechnic Institute of Brooklyn; (1957).	Hr, Kk, Pv, -20 to 150 °C.	B

79GAU:	Gautam, R.; "Computation of physical and chemical Equilibria"; Dissertation, University of Pennsylvania, Philadelphia, (1979). -----	Zc, Kk, 20 °C	C
79GAU/SEI:	Gautam, R.; Seider, W. D.; "Computation of phase and chemical equilibrium. Part I. Local and constrained minima in Gibbs free energy"; AIChE J.; *25, 991 (1979) -----	Zc, Kk, 20 °C	C
80GIG:	Giggenbach, W. F.; "Geothermal gas equilibria"; Geochim. Cosmochim. Acta; *44, 2021 (1980) -----	Zc, Kk, Pv, 100-350 °C.	D
65GIN:	Ginzburg, D. M.; "The System $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O}$ "; J. Appl. Chem. USSR (Eng. Trans.); *38, 2157 (1965) -----	Kk, Pt, Pv, 55-90 °C.	B
72GIN/MAR:	Ginzburg, D. M.; Markel', S. A.; Detinich, L. P.; "The system $\text{NaCl-NH}_4\text{Cl-NH}_3\text{-CO}_2\text{-H}_2\text{O}$ at p = 1 atm"; J. Appl. Chem. USSR (Eng. Trans.); *45, 1762 (1972) -----	Kk, Pv 60-105 °C.	B
65GIN/PIK:	Ginzburg, D. M.; Pikulina, N. S.; Litvin, V. P.; "The system $\text{NH}_3\text{-H}_2\text{S-H}_2\text{O}$ "; J. Appl. Chem. USSR (Eng. Trans.); *38, 2071 (1965) -----	Pv, 64-90 °C.	C
66GIN/PIK:	Ginzburg, D. M.; Pikulina, N. S.; Litvin, V. P.; "The system $\text{NH}_3\text{-H}_2\text{S-H}_2\text{O}$ at 600 mm"; J. Appl. Chem. USSR (Eng. Trans.); *39, 2225 (1966) -----	Kk, Pv, 59-84 °C.	C
67GIO:	Gioia, F.; "Assorbimento simultaneo de H_2S e CO_2 in soluzioni tampone di carbonato e bicarbonato di potassio"; Chim. Ind.; *49, 1287 (1967) -----	Kk, 20 °C.	A
36GME:	Gmelins Handbuch der Anorganischen Chemie; pp. 327-362, Verlag Chemie, Berlin, Germany (1936) -----	Ze, Zr, Kk, Px, Pv, -15 to 197 °C	B

62GOL/ALT:	Goldberg, N. A.; Altschuler, L. N.; "Microscopic kinetics and mechanism for the synthesis of urea from ammonia and carbon dioxide"; Khim. Promst. (Moscow); *9, 638 (1962)	Zc, Kk, 165-195 °C.	B
64GOL/ALT:	Goldberg, N. A.; Altschuler, L. N.; "Macroscopic kinetics and mechanism for the synthesis of urea from ammonia and carbon dioxide"; Khim. Promst. (Moscow); *1, 54 (1964)	Zc, Kk, 170-210 °C.	B
58GOL/MAL:	Golutvin, Yu. M.; Malysheva, T. V.; Skorobogatova, V. I.; "The solubility of hydrogen sulfide and carbon dioxide in water and aqueous solutions of ammonia and phenol"; Akad. Nauk SSSR. Sibir. Otdel. Izvest. Novos.; *8, 83 (1958)	Kk, Pv 20 °C.	B,C
72GOR/GOR:	Gorlovskii, D. M.; Gorbushenkov, V. A.; Kucheryavyi, V. I.; "Experimental values of the equilibrium degree of conversions of carbon dioxide into urea by the Bazarov reaction"; J. Appl. Chem. USSR (Eng. Trans.); *45, 1596 (1972)	Kk, Pv, 165-230 °C, L = 2.9 to 6.3, W = 0 to 1.6.	B
71GOR/KOS:	Gorlovskii, D. M.; Koshcherenkov, N. N.; Kucheryavyi, V. I.; "Equilibrium pressure in the system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ "; J. Appl. Chem. USSR (Eng. Trans.); *44, 2478 (1971)	Kk, Pv, 220-230 °C, L = 2 to 5, W = 0 to 15.	B
73GOR/KOS:	Gorlovskii, D. M.; Koshcherenkov, N. N.; Kucheryavyi, V. I.; "Equilibrium pressure in the system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ "; J. Appl. Chem. USSR (Eng. Trans.); *46, 2618 (1973)	Kk, Pv 160-220 °C, L = 0.5 to 5.0, W = 0.	B
73GOR/KOS2:	Gorlovskii, D. M.; Koshcherenkov, N. N.; Kucheryavyi, V. I.; "Influence of excess carbon dioxide on equilibrium in the Bazarov reaction"; J. Appl. Chem. USSR (Eng. Trans.); *46, 2780 (1973)	Kk, Pv 160-220 °C, L = 0.5 to 5.0, W = 0.	B

76GOR/KOS:	Gorlovskii, D. M.; Koscherenkov, N. N.; Kucheryavii, V. I.; "Diagrams of the thermodynamics properties of the system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ during the synthesis of urea"; J. Appl. Chem. USSR (Eng. Trans.); *49, 780 (1976)	Kk, Pv, 150-250 °C, L = 4, W = 0.	B
78GOR/KOS:	Gorlovskii, D. M.; Koscherenkov, N. N.; Kucheryavii, V. I.; "Interpolation formulas for calculation of the pressure in the system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ with the synthesis of urea"; J. Appl. Chem. USSR (Eng. Trans.); *51, 471 (1978)	Zc, Kk, Pv, 160-220 °C, L = 2.5 to 5.0, W = 0 to 1.5.	B
74GOR/KUC:	Gorlovskii, D. M.; Kucheryavii, V. I.; "Weight ratio of the phases in the equilibrium liquid-gas system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ during synthesis of urea"; J. Appl. Chem. USSR (Eng. Trans.); *47, 205 (1974)	Zc, Kk, Pv, 160-200 °C, L = 2 to 5, W = 0 to 1.5.	B
74GOR/KUC2:	Gorlovskii, D. M.; Kucheryavii, V. I.; Koscherenkov, N. N.; "Equation for the dependence of the equilibrium urea yield by the Bazarov reaction on temperature, composition, and reactor charge density"; J. Appl. Chem. USSR (Eng. Trans.); *47, 230 (1974)	Zc, Kk, Pv, 160-200 °C, L = 2 to 5, W = 0 to 1.5.	B
44GUY/PIE:	Guyver, A.; Piechowicz, T.; "Losungleichgewichte in wässrigen systemen. Das System $\text{CO}_2\text{-NH}_3\text{-H}_2\text{O}$ bei 20° and 50°"; Helv. Chim. Acta; *27, 858 (1944)	Pv, 20-50 °C, 0 to 35 percent CO_2 , 0 to 35 percent NH_3 .	B
79HAL/DRE:	Hales, J. M.; Drewes, D. R.; "Solubility of ammonia in water at low concentrations"; Atmos. Environ.; *13, 1133 (1979)	Kk, Pv, 277-296 K, H_2SO_4 also present.	B
79HEG/MOL:	Hegner, B.; Molzahn, M.; "Simulation of mass transfer controlled countercurrent separation processes accompanied by chemical reactions"; Inst. Chem. Eng. Symp. Ser. No. 56; pp. 81-100 (1979)	Zc, Kk, 60-90 °C	B
28HOL:	Hollings, H.; "Some application of chemistry in gas making"; Trans. Inst. Gas. Eng.; (1927-28), pp. 495-545	Kk, 50-90 °F, 3.5 percent CO_2 .	B

42HOR/OGA:	Hori, S.; Ogami, N.; "The synthesis of urea from ammonia and carbon dioxide"; Koatsu Gasu Kyokaishi; #6, 256 (1942)	Kk, 119-220 °C.	B
61IKE:	Ikeno, S.; "Granulation of fertilizer from ammonium carbamate solution"; Kogyo Kagaku Zasshi; #64, 627 (1971)	Pv, 15-40 °C.	B
72INO/KAN:	Inoue, S.; Kanai, K.; Otsuka, E.; "Equilibrium of urea synthesis. I"; Bull. Chem. Soc. Japan; #45, 1339 (1972)	Kk, Pv, 160-220 °C, L = 0 to 8, W = 0 to 1.2.	B
72INO/KAN2:	Inoue, S.; Kanai, K.; Otsuka, E.; "Equilibrium of urea synthesis. II"; Bull. Chem. Soc. Japan; #45, 1616 (1972)	Kk, Pv, 160-220 °C, L = 0 to 8, W = 0 to 1.2.	B
77ISA/OTT:	Isaacs, E. E.; Otto, F. D.; Mather, A. E.; "The solubility of mixtures of carbon dioxide and hydrogen sulphide in an aqueous DIPA solution"; Can. J. Chem. Eng.; #55, 210 (1977)	Pv, 40-100 °C, in sols of di-isopropanol-amine, p(CO ₂) = 2 to 5991 kPa, p(H ₂ S) = 1.3 to 4126 kPa.	A
80ISA/OTT:	Isaacs, E. E.; Otto, F. D.; Mather, A. E.; "Solubility of mixtures of H ₂ S and CO ₂ in amonoethanolamine solution at low partial pressures"; J. Chem. Eng. Data; #25, 118 (1980)	Pv, 100 °C. in monoethanolamine, p(CO ₂) = 7 to 1360 Pa, p(H ₂ S) = 0 to 3920 Pa.	A
29JAN:	Janecke, E.; "Über das System H ₂ O, CO ₂ und NH ₃ "; Z. Elektrochem.; #35, 716 (1929)	Pv, 0-90 °C.	B
30JAN:	Janecke, E.; "Über das System" H ₂ O-CO ₂ -NH ₃ , Fortsetzung"; Z. Elektrochem.; #36, 645 (1930)	Pv, Px, -78 to 132 °C	B
59JON/FRO:	Jones, J. H.; Froning, H. R.; Claytor, Jr., E. E.; "Solubility of acidic gases in aqueous monoethanolamine"; J. Chem. Eng. Data; #4, 85 (1959)	Pv, 40-140 °C, in monoethanolamine.	A

71KAA:	Kaasenbrood, P. J. C.; "The urea stripping process: The technical manufacture of urea, with carbon dioxide used both as reactant and as stripping gas"; Proc. Fourth European Symposium held in Brussels, Sept. 9-11, 1968, Supplement to Chemical Engineering Science, Pergamon Press, Oxford (1971)	Zr, Px, 150-190 °C	B
77KAA/CHE:	Kaasenbrood, P. J. C.; "Stripping technology, phase equilibria and thermodynamics of the stamicarbon urea stripping process"; Chermin, H. A. G.; Proc. Fert. Soc. 3, (1977)	Zr; Kk, Pv, 155-190 °C, 0 to 35 mass percent CO ₂ , 0 to 35 mass percent NH ₃ .	B
74KAT/VAI:	Katkovskaya, K. Ya.; Vaineikis, A. A.; Dubrovskii, I. Ya.; "Influence of temperature on pH in the system H ₂ O-CO ₂ -NH ₃ "; Teploenergetika (Moscow); *7, 8 (1974)	Ke, 25-250 °C.	B
51KAW:	Kawasumi, S.; "Equilibrium of the CO ₂ -NH ₃ -Urea-H ₂ O system under high temperature and pressure. I. Equilibrium pressure of reaction in urea-synthesis from ammonia and carbon dioxide"; Bull. Chem. Soc. Japan; *24, 148 (1951)	Kk, Pv, 130-170 °C, 30 to 185 atm.	B
52KAW:	Kawasumi, K.; "Equilibrium of the CO ₂ -NH ₃ -Urea-H ₂ O system under high temperature and pressure. II. Liquid-vapor equilibrium in the loading mole ratio of NH ₃ to CO ₂ "; Bull. Chem. Soc. Japan; *25, 227 (1952)	Pv, Xd, 140-180 °C, L = 2.	B
53KAW:	Kawasumi, S.; "Equilibrium of the CO ₂ -NH ₃ -Urea-H ₂ O system under high temperature and pressure. III. Effect of water added on liquid-vapor equilibrium"; Bull. Chem. Soc. Japan; *26, 218 (1953)	Kk, Pv, 130-190 °C, L = 1.5 to 4.0.	B
53KAW2:	Kawasumi, S.; "Equilibrium of the CO ₂ -NH ₃ -urea-H ₂ O system under high temperature and pressure. IV. Effect of loading NH ₃ -CO ₂ mole ratio on equilibrium pressure and vapor composition"; Bull. Chem. Soc. Japan; *26, 222 (1953)	Kk, Pv, 130-190 °C, L = 1.5 to 4.0.	B

54KAW:	Kawasumi, S.; "Equilibrium of the CO ₂ -NH ₃ -urea-H ₂ O system under high temperature and pressure. V. Liquid-vapor equilibrium in the presence of excess ammonia or carbon dioxide"; Bull. Chem. Soc. Japan; *27, 254 (1954)	Kk, Pv, 160 °C, L = 1.5 to 3.3.	B
76KEN/EIS:	Kent, R. L.; Eisenberg, B.; "Better data for amine treating"; Hydrocarbon Process.; *55, 87 (1976)	Zc, Kk, Pv, 0-140 °C,	A
73KHE/SAM:	Khetchikov, L. N.; Samoilovich, L. A.; "Correlation of experimental data (p-ρ-T-x) in salt-water and salt-gas-water systems"; Exp. Miner. Petrogr.; pp. 112-121 (1973)	Pv, Xd, 160-400 °C,	B
51KIN:	Kinoshita, H.; "Equilibrium of urea-water system. II"; Rev. Phys. Chem. Japan; *21, 16 (1951)	Pv, 150-210 °C, 32 to 306 atm.	B
51KIY/KIN:	Kiyama, R.; Kinoshita, H.; "Equilibrium of urea-water system. I"; Rev. Phys. Chem. Japan; *21, 9 (1951)	Pv, 130-240 °C, 0 to 500 atm.	B
71KON/BOL:	Koneczny, H.; Bolinski, L.; "Nature of the equilibrium of the sodium process at 25 °C"; Przem. Chem.; *50, 657 (1971)	Kk, Pv, 25 °C, NaCl present.	B
74KON/TRY:	Koneczny, H.; Trypuc, M.; "Influence of temperature on the equilibrium pressure over the system NaNO ₃ -NH ₃ -CO ₂ -H ₂ O"; Chem. Stosow.; *18, 15 (1974)	Kk, 20-40 °C.	B
79KOT/DUK:	Kotula, E.; Dukowicz, J.; Wojcicka, K.; "Determination of the specific heat of ammonia carbon dioxide-water-urea solutions"; Pr. Nauk. Inst. Technol. Nieorg. Nazow Miner. Politech. Wroclaw.; *15, 153 (1979); [CA 92-136312k].	Q1, 303-393 K, 2.5 MPa, 0-60 wt% urea, 1-40 wt% NH ₃ , 0-44 wt% CO ₂ .	B
80KOT/KOW:	Kotula, E.; Kowalik, W.; "Determination of the equilibrium pressure of ammonia in the system NH ₃ -CO ₂ -H ₂ O-urea up to pressures of 2.5 MPa"; Chem. Stosow.; *24, 561 (1980)	Pv, 80-1120 °C, 1.8 to 2.5 MPa	B

75KOU:	Koubsky, P.; "Vodivost vodneho rotoku systemu $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O}$ "; Chem. Prum.; *25, 291 (1975)	Zc, Kk, Pv, 25 °C.	B
74KOU/HLA:	Koubsky, P.; Hladky, V.; "Zpracovani polynovych kondensatu pri vyrobe svitiplynu-reseni systemu $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O}$ "; Plyn.; *54, 337 (1974)	Zc, Kk, Pv, 25-80 °C.	B
75KOU/HLA:	Koubsky, P.; Hladky, V.; "Vypocet rovnovah v systemu $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O}$ "; Chem. Prum.; *25, 234 (1975)	Zc, Kk, Pv, 25-80 °C.	B
76KOU/HLA:	Koubsky, P.; Hladky, V.; "Equilibrium calculations in the system $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O}$ "; Int. Chem. Eng.; *16, 392 (1976)	Zc, Kk, Pv, 25-80 °C	B
63KOZ/KAR:	Kozin, V. M.; Karpukhin, A. M.; Momot, M. V.; Volkov, B. V.; "Study of the equilibrium of ammonia and carbon dioxide over boric acid-glycerol aqueous solutions"; Khim. Prom. Nauk. Tekhn. Zb.; pp. 10-14 (1963)	Kk, Pv, 50-150 °C, boric acid and glycerol are also present.	B
22KRA/GAD:	Krase, N. W.; Gaddy, V. L.; "Synthesis of urea from ammonia and carbon dioxide"; Ind. Eng. Chem.; *14, 611 (1922)	Kk, Pv, 70-200 °C.	B
30KRA/GAD:	Krase, H. J.; Gaddy, V. L.; "Equilibria in the ammonium carbamate-urea-water ammonia system"; J. Am. Chem. Soc.; *52, 3088 (1930)	Kk, 155 °C.	B
38KRA:	Krasil'shchikov, A. I.; "The physical chemical basis for the synthesis of urea"; Usp. Khim.; *7, 1042 (1938)	Zc, Kk, 135-200 °C.	B
49KRE/HOF:	van Krevelen, D. W.; Hoftijzer, P. J.; Huntjens, F. J.; "Composition and vapour pressures of aqueous solution of ammonia, carbon dioxide and hydrogen sulphide"; Rec. Trav. Chem. *68, 191 (1949)	Kk, Pv, 20-60 °C.	A,B,C,D

69KUC/GOR:	Kucheryavii, V. I.; Gorbushenkov, V. A.; "Reaction kinetics for the synthesis of urea using Bazarov's process"; Dokl. Akad. Nauk SSSR; *188, 868 (1969)	Pv, Xd, 150-210 °C, L = 2 to 5, W = 0 to 1.0.	B
69KUC/GOR2:	Kucheryavii, V. I.; Gorlovskii, D. M.; "Reaction equilibria for the synthesis of urea from ammonia and carbon dioxide"; Khim. Promst. (Moscow); *11, 836 (1969)	Pv, Xd, 150-210 °C, L = 2 to 5, W = 0 to 1.0.	B
69KUC/GOR3:	Kucheryavii, V. I.; Gorlovskii, D. M.; "The physical chemical behavior of vapor liquid equilibria in the system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ with parameters for the synthesis of urea"; Dokl. Akad. Nauk SSSR; *186, 891 (1969)	Zc, Kk, 170-210 °C, L = 2 to 6, W = 0 to 1.	B
69KUC/GOR4:	Kucheryavii, V. I.; Gorlovskii, D. M.; Konkina, T. N.; "The steps of the equilibria for the conversion of carbon dioxide to urea"; Khim. Promst. (Moscow); *3, 200 (1969)	Zc, Kk, 100-220 °C, L = 0 to 7.6, W = 0 to 7.	B
69KUC/GOR5:	Kucheryavii, V. I.; Gorlovskii, D. M.; Polyakov, E. V.; "Dependence of equilibrium pressure in the system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ (gas-liquid) on the temperature, autoclave charge density and $\text{NH}_3\text{:CO}_2$ ratio in the original component mixture"; J. Appl. Chem. USSR (Eng. Trans.); *42, 591 (1969)	Pv, 133-241 °C, L = 4.	B
69KUC/GOR6:	Kucheryavii, V. I.; Gorlovskii, D. M.; Konkina, T. N.; "Volume fractions of liquid and gas phases in the equilibrium system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ at high temperatures and pressures"; J. Appl. Chem. USSR (Eng. Trans.); *42, 1020 (1969)	Pv, 169-190 °C, L = 2 to 3.3, W = 0 to 1.0.	B
69KUC/GOR7:	Kucheryavii, V. I.; Gorbushenkov, V. A.; "Influence of excess water and ammonia on the rate of synthesis of urea from ammonia and carbon dioxide"; J. Appl. Chem. USSR (Eng. Trans.); *42, 2459 (1969)	Kk, 190 °C, L = 4, W = 0.5.	B

69KUC/GOR8:	Kucheryavii, V. I.; Gorbushenkov, V. A.; "Kinetics of urea synthesis from ammonia and carbon dioxide"; J. Appl. Chem. USSR (Eng. Trans.); #42, 2293 (1969)	Kk, Pv, 165-210 °C, L = 4.	B
70KUC/GOR:	Kucheryavii, V. I.; Gorbushenkov, V. A.; "Dependence of the rate of urea synthesis from ammonia and carbon dioxide on the density of gas-liquid mixture"; J. Appl. Chem. USSR (Eng. Trans.); #43, 218 (1970)	Pv, Xd, 140-190 °C, L = 2 to 5, W = 0.2 to 1.0.	B
70KUC/GOR2:	Kucheryavii, V. I.; Gorlovskii, D. M.; "Densities of coexisting phases in the equilibrium gas-liquid system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ at high temperatures and pressures"; J. Appl. Chem. USSR (Eng. Trans.); #43, 1693 (1970)	Pv, Xd, 140-190 °C, L = 2 to 5, W = 0.2 to 1.0.	B
72KUC/GOR:	Kucheryavii, V. I.; Gorlovskii, D. M.; Koshcharenkov, N. N.; "Influence of excess urea on equilibrium in the Bazarov reaction"; J. Appl. Chem. USSR (Eng. Trans.); #45, 70 (1972)	Kk, Pv, 165-200 °C, L = 2.5 to 5.0, W = 0.0 to 0.8.	B
75KUC/GOR:	Kucheryavii, V. I.; Gorlovskii, D. M.; Koshcharenkov, N. N.; "Phase diagram, critical parameters, and azeotropy of the system $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ "; J. Appl. Chem. USSR (Eng. Trans.); #48, 320 (1975)	Pt, Pv, 150-240 °C, L = 4.	B
68KUC/ZIN:	Kucheryavii, V. I.; Zinov'ev, G. N.; Koshcharenkov, N. N.; "Liquid-vapor equilibrium in the system ammonia-carbon dioxide-urea-water at pressures up to 50 atm and temperatures in the range of 100-160°"; J. Appl. Chem. USSR (Eng. Transl.); #41, 795 (1968)	Kk, Pv, 100-160 °C.	B
69KUC/ZIN:	Kucheryavii, V. I.; Zinov'ev, G. N.; "Calculation of the equilibrium distribution of NH_3 , CO_2 , and H_2O between fluids due to both transport and pressure in the systems $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O}$ and $\text{NH}_3\text{-CO}_2\text{-CO(NH}_2)_2\text{-H}_2\text{O}$ "; Khim. Promst. (Moscow); #5, 354 (1969)	Zc, Pv, 100-150 °C, 20 to 50 atm.	B

70KUR/ILI:	Kurshev, I.; Ilieva, E.; Shishkov, D.; "A new modification of a dynamic method for determining equilibrium partial pressures of water, ammonia, and carbon dioxide"; God Nauchnoizsled. Inst. Khim. Promst.; *8, 81 (1971) -----	Pv, t = ?	B
70KUR/ILI2:	Kurshev, I.; Ilieva, E.; Shishkov, D.; "A new modification of a dynamic method for determining equilibrium partial pressure of water, ammonia, and carbon dioxide"; Khim. Ind. (Sofia); *10, 440 (1970) -----	Pv, t = ?	B
71KUR/ILI:	Kurshev, I.; Ilieva, E.; "Determination of equilibrium partial pressures of ammonia and carbon dioxide in the ammonia-carbon dioxide-ammonium nitrate-water system"; God. Nauchnoizsled. Inst. Khim. Promst.; *8, 87 (1971) -----	Pv, 20-40 °C, L = 1.7 to 4.5.	B
79LAB/EKG:	Labedeva, G. N.; Ekgauz, V. I.; Anikina, T. G.; Zabello, L. A.; "Partial pressures of ammonia, hydrogen sulfide, and hydrogen cyanide over their aqueous solution"; J. Appl. Chem. USSR (Eng. Trans.); *7, 1424 (1979) -----	Pv, 20-40 °C, HCN also present.	D
76LAW/GAR:	Lawson, J. D.; Garst, A. W.; "Gas sweetening data: equilibrium solubility of hydrogen sulfide and carbon dioxide in aqueous monoethanolamine and aqueous diethanolamine solutions"; J. Chem. Eng. Data; *21, 20 (1976) -----	Kk, Pv, 100-300 °F, in monoethanolamine and diethanolamine sols.	A
74LEE/OTT:	Lee, J. I.; Otto, F. D.; Mather, A. E.; "The solubility of mixtures of carbon dioxide and hydrogen sulphide in aqueous diethanolamine solutions"; Can. J. Chem. Eng.; *52, 125 (1974) -----	Kk, Pv, 40 °C and 100 °C, in monoethanolamine sols, p(CO ₂) = 0.1 to 1000 psia, p(H ₂ S) = 0.3 to 650 psia.	A
74LEE/OTT2:	Lee, J. I.; Otto, F. D.; Mather, A. E.; "The solubility of H ₂ S and CO ₂ in aqueous monoethanolamine solutions"; Can. J. Chem. Eng.; *52, 803 (1974) -----	Kk, Pv, 40 °C and 100 °C, in monoethanolamine sols, p(CO ₂) = 0.1 to 1000 psia, p(H ₂ S) = 0.3 to 650 psia.	A

75LEE/OTT:	Lee, J. I.; Otto, F. D.; Mather, A. E.; "Solubility of mixtures of carbon dioxide and hydrogen sulfide in 5.0N monoethanolamine solution"; J. Chem. Eng. Data; *20, 161 (1975)	Kk, Pv, 40 °C and 100 °C, in 5.0 N monoethanolamine sols, p(CO ₂) = 0.1 to 810 psia, p(H ₂ S) = 0.1 to 510 psia.	A
76LEE/OTT:	Lee, J. I.; Otto, F. D.; Mather, A. E.; "The measurement and prediction of the solubility of mixtures of carbon dioxide and hydrogen sulphide in a 2.5 N monoethanolamine solution"; Can. J. Chem. Eng.; *54, 214 (1976)	Kk, Pv, 40 °C and 100 °C, in 2.5 N monoethanolamine sols, p(CO ₂) = 0.7 to 5630 kPa, p(H ₂ S) = 0.7 to 3780 kPa.	A
50LEI/SHE:	Leibush, A. G.; Sheerson, A. L.; "Solubility of hydrogen sulfide in mixtures containing carbon dioxide and ethanolamine"; Zh. Prikl. Khim. (Leningrad); *23, 145 (1950)	Pv, 25 °C, in ethanolamine sols, p(H ₂ S) = 1.3 to 240 mmHg, p(CO ₂) = 0 to 75 mmHg.	A
75LEM:	Lemkowitz, S. M.; "Phase and corrosion studies of the ammonia-carbon dioxide-water system at the conditions of the hot gas recirculation process for the synthesis of urea"; Dissertation, Delft, The Netherlands (1975).	Px	B
72LEM/COO:	Lemkowitz, S. M.; de Cooker, M. G. R. T.; van den Berg, P. J.; "Some fundamental aspects of urea technology"; Proc. Fert. Soc. No. 131; pp. 3 - 115 (1972)	Zc, Kk, Pv, Pt, Pt, Pv, 130-200 °C	B
73LEM/COO:	Lemkowitz, S. M.; de Cooker, M. G. R. T.; van den Berg, P. J.; "An empirical thermodynamic model for the ammonia water-carbon dioxide system at urea synthesis conditions"; J. Appl. Chem. Biotechnol.; *23, 63 (1973)	Zc, Kk, Pt, Pv, 140-200 °C, 50 to 250 atm, L = 2.4 to 6, W F 0.5.	B
77LEM/DIE:	Lemkowitz, S. M.; Diepen, G. A. M.; van den Berg, P. J.; "A phase model for the gas-liquid equilibria in the ammonia-carbon dioxide water-urea system in chemical equilibrium at urea synthesis conditions. I. Theory"; J. Appl. Chem. Biotechnol; *27, 327 (1977)	Zc, Kk, Pv, Pt, 140-220 °C, 10 to 300 atm.	B

71LEM/GOE:	Lemkowitz, S. M.; Goedegebuur, J.; van den Berg, P. J.; "Bubble-point measurements in the ammonia-carbon dioxide system"; J. Appl. Chem. Biotechnol.; *21, 229 (1971)	Pt, Pv, 130-185 °C, 30 to 175 atm.	B
78LEM/ERP:	Lemkowitz, S. M.; van Erp, J. C. V.; van den Berg, P. J.; "Dew-point measurements of ammonia-water-carbon dioxide mixtures at urea syntheses conditions"; Chem. Chem. Eng.; *3, 83 (1978)	Kk, Pt, Pv, 140-180 °C, 30 to 200 atm, 0 to 15 mole percent H ₂ O.	B
80LEM/ERP:	Lemkowitz, S. M.; van Erp, J. C.; Rekers, D. M.; van den Berg, P. J.; "Phase equilibria in ammonia-carbon dioxide systems at and above urea synthesis conditions"; J. Chem. Techn. Biotechnol.; *30, 85 (1980)	Kk, Pt, Pv, 150-300 °C, L = 2.	B
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66MOS:	Mostofin, A. A.; "Calculation of the effects of pH and specific electrolytes in gaseous solutions of NH ₃ and CO ₂ "; Vodn. Rezhim. Khim. Kont. Paros. Ustand. Moscow; *2, 178 (1966)	Zc, Kk, 25 °C, pH = 4-12.	B
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500TS/TAN:	Otsuka, E.; Tanimoto, K.; Yakabe, M.; "Method of recovery of excess ammonia in urea synthesis"; Kogyo Kazaku Zasshi; *62, 29 (1959)	Kk, 64-129 °C.	B
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600TS/TAK:	Otsuka, E.; Takahashi, T.; "Equilibrium in the system Na ₂ CO ₃ -NH ₃ -CO ₂ -H ₂ O"; Kogyo Kagaku Zasshi; *63, 1218 (1960)	Pv, 71-161 °C.	B
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83OWE/CUN:	Owens, J. L.; Cunningham, J. R.; Wilson, G. M.; 'Vapor liquid equilibria for sour water systems at high temperatures', Research Report RR-65; Gas Producer's Association, Tulsa, OK (1983). -----	Pv, 200-500 °F.	B, C, D
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82PAW/NEW:	Pawlikowski, E. M.; Newman, J.; Prausnitz, J. M.; "Phase equilibria for aqueous solutions of ammonia and carbon dioxide"; Ind. Eng. Chem. Proc. Des. Dev.; *21, 764 (1982) -----	Kk, Pv, 100 and 150 °C, measurements also done in aqueous KNO ₃ and K ₂ SO ₄ .	B
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Zc, Vt, 30-95 °C. B
- 80SA: Szarawara, J.; Piotrowski, J.; Zolotajkin, M.; "Equilibrium pressures over the system $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O-urea}$ "; Zesz. Nauk. Poltech. Slask. Chem.; *677, 43 (1980)

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80TSE:	Tseitlin, M. A.; "On the calculation of vapor-liquid equilibria for the system $\text{NH}_3\text{-CO}_2\text{-H}_2\text{O-NaCl}$ "; Zh. Prikl. Khim. (Leningrad); *53, 2343 (1980)	Zc, Pv, 60 °C.	B

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63WIC:	Wicar, S.; "Calculation of vapour-liquid equilibrium in the system urea, ammonia, carbon dioxide, and water"; Brit. Chem. Eng.; *8, 818 (1963)	Zc, Kk, Pv, 60-160 °C.	B
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