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\(\dot{n} = 0.985 \cdot 0.7219 \cdot \frac{0.00456 \cdot 98836}{\sqrt{1 \cdot 0.0287805 \cdot 8.314472 \cdot 378.15}}\)

\(\dot{n} = 33.690\ \text{mol/s}\)


\(\$ 1065.644\) Vacuum-decay leak rate.

This section describes how to calculate the leak rate of a vacuum-decay leak verification, which is described in \$1065.345(e). Use Eq. 1065.644–1 to calculate the leak rate, \(\dot{n}_{\text{leak}}\), and compare it to the criterion specified in \$1065.345(e).

\[
\dot{n}_{\text{leak}} = \frac{V_{\text{vac}}}{R} \left( \frac{p_2 - p_1}{T_2 - T_1} \right) \quad \text{Eq. 1065.644-1}
\]

Where:

\(V_{\text{vac}}\) = geometric volume of the vacuum-side of the sampling system.

\(R\) = molar gas constant.

\(p_2\) = Vacuum-side absolute pressure at time \(t_2\).

\(T_2\) = Vacuum-side absolute temperature at time \(t_2\).

\(p_1\) = Vacuum-side absolute pressure at time \(t_1\).

\(T_1\) = Vacuum-side absolute temperature at time \(t_1\).

\(t_2\) = time at completion of vacuum-decay leak verification test.

\(t_1\) = time at start of vacuum-decay leak verification test.

Example:

\(V_{\text{vac}} = 2.0000\ \text{L} = 0.00200\ \text{m}^3\)

\(R = 8.314472\ \text{J/(mol·K)}\)

\(p_2 = 50.600\ \text{kPa} = 50600\ \text{Pa}\)

\(T_2 = 293.15\ \text{K}\)

\(p_1 = 25.300\ \text{kPa} = 25300\ \text{Pa}\)

\(T_1 = 293.15\ \text{K}\)

\(t_2 = 10:57:35\ \text{AM}\)

\(t_1 = 10:56:25\ \text{AM}\)

\[
\dot{n}_{\text{leak}} = \frac{0.0002}{8.314472} \cdot \frac{50600 - 25300}{293.15 - 293.15} \quad \text{Eq. 1065.644-1}
\]

\[
\dot{n}_{\text{leak}} = \frac{0.00200}{8.314472} \cdot \frac{86.304}{70}
\]

\(\dot{n}_{\text{leak}} = 0.00030\ \text{mol/s}\)

[73 FR 37327, June 30, 2008]

\(\$ 1065.645\) Amount of water in an ideal gas.

This section describes how to determine the amount of water in an ideal gas, which you need for various performance verifications and emission calculations. Use the equation for the vapor pressure of water in paragraph (a) of this section or another appropriate equation and, depending on...