

§ 1065.655

40 CFR Ch. I (7-1-11 Edition)

test intervals based on the ratio of mass rate to power as described in paragraph (b)(2) of this section:

$$e_{\text{composite}} = \frac{\sum_{i=1}^N WF_i \cdot \bar{m}_i}{\sum_{i=1}^N WF_i \cdot \bar{P}_i} \quad \text{Eq. 1065.650-19}$$

Where:

- i = test interval number.
- N = number of test intervals.
- WF = weighting factor for the test interval as defined in the standard-setting part.
- \bar{m} = mean steady-state mass rate of emissions over the test interval as determined in paragraph (e) of this section.
- \bar{P} is the mean steady-state power over the test interval as described in paragraph (e) of this section.

Example:

- $N = 2$
- $WF_1 = 0.85$
- $WF_2 = 0.15$
- $\bar{m}_1 = 2.25842 \text{ g/hr}$
- $\bar{m}_2 = 0.063443 \text{ g/hr}$
- $\bar{P}_1 = 4.5383 \text{ kW}$
- $\bar{P}_2 = 0.0 \text{ kW}$

$$e_{\text{NO}_x \text{ composite}} = \frac{(0.85 \cdot 2.25842) + (0.15 \cdot 0.063443)}{(0.85 \cdot 4.5383) + (0.15 \cdot 0.0)}$$

$e_{\text{NO}_x \text{ composite}} = 0.5001 \text{ g/kW}\cdot\text{hr}$

(h) *Rounding.* Round the final brake-specific emission values to be compared to the applicable standard only after all calculations are complete (including any drift correction, applicable deterioration factors, adjustment factors, and allowances) and the result is in g/(kW·hr) or units equivalent to the units of the standard, such as g/(hp·hr). See the definition of “Round” in §1065.1001.

[73 FR 37328, June 30, 2008, as amended at 73 FR 59332, Oct. 8, 2008; 75 FR 23048, Apr. 30, 2010]

§1065.655 Chemical balances of fuel, intake air, and exhaust.

(a) *General.* Chemical balances of fuel, intake air, and exhaust may be used to calculate flows, the amount of water in their flows, and the wet concentration of constituents in their flows. With one flow rate of either fuel, intake air, or exhaust, you may use chemical balances to determine the flows of the other two. For example,

you may use chemical balances along with either intake air or fuel flow to determine raw exhaust flow.

(b) *Procedures that require chemical balances.* We require chemical balances when you determine the following:

(1) A value proportional to total work, \bar{W} , when you choose to determine brake-specific emissions as described in §1065.650(e).

(2) The amount of water in a raw or diluted exhaust flow, $x_{\text{H}_2\text{Oexh}}$, when you do not measure the amount of water to correct for the amount of water removed by a sampling system. Correct for removed water according to §1065.659(c)(2).

(3) The flow-weighted mean fraction of dilution air in diluted exhaust, $x_{\text{dil/exh}}$, when you do not measure dilution air flow to correct for background emissions as described in §1065.667(c). Note that if you use chemical balances for this purpose, you are assuming that your exhaust is stoichiometric, even if it is not.

(c) *Chemical balance procedure.* The calculations for a chemical balance involve a system of equations that require iteration. We recommend using a computer to solve this system of equations. You must guess the initial values of up to three quantities: The amount of water in the measured flow, $x_{\text{H}_2\text{Oexh}}$, fraction of dilution air in diluted exhaust, $x_{\text{dil/exh}}$, and the amount of products on a C_1 basis per dry mole of dry measured flow, x_{Ccombdry} . You may use time-weighted mean values of combustion air humidity and dilution air humidity in the chemical balance; as long as your combustion air and dilution air humidities remain within tolerances of ± 0.0025 mol/mol of their respective mean values over the test interval. For each emission concentration, x , and amount of water, $x_{\text{H}_2\text{Oexh}}$, you must determine their completely dry concentrations, x_{dry} and $x_{\text{H}_2\text{Oexhdry}}$. You must also use your fuel's atomic hydrogen-to-carbon ratio, α , oxygen-to-carbon ratio, β , sulfur-to-carbon ratio, γ , and nitrogen-to-carbon ratio, δ . You may measure α , β , γ , and δ or you may use default values for a given fuel as described in §1065.655(d). Use the following steps to complete a chemical balance:

(1) Convert your measured concentrations such as, $x_{\text{CO}_2\text{meas}}$, x_{NOmeas} , and $x_{\text{H}_2\text{Oint}}$, to dry concentrations by dividing them by one minus the amount of water present during their respective measurements; for example: $x_{\text{H}_2\text{OxCO}_2\text{meas}}$, $x_{\text{H}_2\text{OxNOmeas}}$, and $x_{\text{H}_2\text{Oint}}$. If the amount of water present during a "wet" measurement is the same as the unknown amount of water in the exhaust flow, $x_{\text{H}_2\text{Oexh}}$, iteratively solve for that value in the system of equations. If you measure only total NO_x and not NO and NO_2 separately, use good engineering judgment to estimate a split in your total NO_x concentration between NO and NO_2 for the chemical balances. For example, if you measure emissions from a stoichiometric spark-ignition engine, you may assume all NO_x is NO. For a compression-ignition engine, you may assume that your molar concentration of NO_x , x_{NOx} , is 75% NO and 25% NO_2 . For NO_2 storage aftertreatment systems, you may assume x_{NOx} is 25% NO and 75% NO_2 . Note that for calculating the mass of NO_x

emissions, you must use the molar mass of NO_2 for the effective molar mass of all NO_x species, regardless of the actual NO_2 fraction of NO_x .

(2) Enter the equations in paragraph (c)(4) of this section into a computer program to iteratively solve for $x_{\text{H}_2\text{Oexh}}$, x_{Ccombdry} , and $x_{\text{dil/exh}}$. Use good engineering judgment to guess initial values for $x_{\text{H}_2\text{Oexh}}$, x_{Ccombdry} , and $x_{\text{dil/exh}}$. We recommend guessing an initial amount of water that is about twice the amount of water in your intake or dilution air. We recommend guessing an initial value of x_{Ccombdry} as the sum of your measured CO_2 , CO, and THC values. We also recommend guessing an initial $x_{\text{dil/exh}}$ between 0.75 and 0.95, such as 0.8. Iterate values in the system of equations until the most recently updated guesses are all within $\pm 1\%$ of their respective most recently calculated values.

(3) Use the following symbols and subscripts in the equations for this paragraph (c):

$x_{\text{dil/exh}}$ = amount of dilution gas or excess air per mole of exhaust.

$x_{\text{H}_2\text{Oexh}}$ = amount of water in exhaust per mole of exhaust.

x_{Ccombdry} = amount of carbon from fuel in the exhaust per mole of dry exhaust.

$x_{\text{H}_2\text{dry}}$ = amount of H_2 in exhaust per amount of dry exhaust.

$K_{\text{H}_2\text{Ogas}}$ = water-gas reaction equilibrium coefficient. You may use 3.5 or calculate your own value using good engineering judgment.

$x_{\text{H}_2\text{Oexhdry}}$ = amount of water in exhaust per dry mole of dry exhaust.

$x_{\text{prod/intdry}}$ = amount of dry stoichiometric products per dry mole of intake air.

$x_{\text{dil/exhdry}}$ = amount of dilution gas and/or excess air per mole of dry exhaust.

$x_{\text{int/exhdry}}$ = amount of intake air required to produce actual combustion products per mole of dry (raw or diluted) exhaust.

$x_{\text{raw/exhdry}}$ = amount of undiluted exhaust, without excess air, per mole of dry (raw or diluted) exhaust.

$x_{\text{O}_2\text{int}}$ = amount of intake air O_2 per mole of intake air.

$x_{\text{CO}_2\text{intdry}}$ = amount of intake air CO_2 per mole of dry intake air. You may use $x_{\text{CO}_2\text{intdry}} = 375 \mu\text{mol/mol}$, but we

recommend measuring the actual concentration in the intake air.
 $x_{\text{H}_2\text{Ointdry}}$ = amount of intake air H₂O per mole of dry intake air.
 $x_{\text{CO}_2\text{int}}$ = amount of intake air CO₂ per mole of intake air.
 $x_{\text{CO}_2\text{dil}}$ = amount of dilution gas CO₂ per mole of dilution gas.
 $x_{\text{CO}_2\text{dildry}}$ = amount of dilution gas CO₂ per mole of dry dilution gas. If you use air as diluent, you may use $x_{\text{CO}_2\text{dildry}} = 375 \mu\text{mol/mol}$, but we recommend measuring the actual concentration in the intake air.
 $x_{\text{H}_2\text{Odildry}}$ = amount of dilution gas H₂O per mole of dry dilution gas.
 $x_{\text{H}_2\text{Odil}}$ = amount of dilution gas H₂O per mole of dilution gas.
 $x_{\text{[emission]meas}}$ = amount of measured emission in the sample at the respective gas analyzer.
 $x_{\text{[emission]dry}}$ = amount of emission per dry mole of dry sample.
 $x_{\text{H}_2\text{O[emission]meas}}$ = amount of water in sample at emission-detection loca-

tion. Measure or estimate these values according to §1065.145(e)(2).
 $x_{\text{H}_2\text{Oint}}$ = amount of water in the intake air, based on a humidity measurement of intake air.
 α = atomic hydrogen-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.
 β = atomic oxygen-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.
 γ = atomic sulfur-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.
 δ = atomic nitrogen-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.
 (4) Use the following equations to iteratively solve for $x_{\text{dil/exh}}$, $x_{\text{H}_2\text{Oexh}}$, and x_{Ccombdry} :

$$x_{\text{dil/exh}} = 1 - \frac{x_{\text{raw/exhdry}}}{1 + x_{\text{H}_2\text{Oexhdry}}} \quad \text{Eq. 1065.655-1}$$

$$x_{\text{H}_2\text{Oexh}} = \frac{x_{\text{H}_2\text{Oexhdry}}}{1 + x_{\text{H}_2\text{Oexhdry}}} \quad \text{Eq. 1065.655-2}$$

$$x_{\text{Ccombdry}} = x_{\text{CO}_2\text{dry}} + x_{\text{COdry}} + x_{\text{THCdry}} - x_{\text{CO}_2\text{dil}} \cdot x_{\text{dil/exhdry}} - x_{\text{CO}_2\text{int}} \cdot x_{\text{int/exhdry}} \quad \text{Eq. 1065.655-3}$$

$$x_{\text{H}_2\text{dry}} = \frac{x_{\text{COdry}} \cdot (x_{\text{H}_2\text{Oexhdry}} - x_{\text{H}_2\text{Odil}} \cdot x_{\text{dil/exhdry}})}{K_{\text{H}_2\text{O-gas}} \cdot (x_{\text{CO}_2\text{dry}} - x_{\text{CO}_2\text{dil}} \cdot x_{\text{dil/exhdry}})} \quad \text{Eq. 1065.655-4}$$

$$x_{\text{H}_2\text{Oexhdry}} = \frac{\alpha}{2} (x_{\text{Ccombdry}} - x_{\text{THCdry}}) + x_{\text{H}_2\text{Odil}} \cdot x_{\text{dil/exhdry}} + x_{\text{H}_2\text{Oint}} \cdot x_{\text{int/exhdry}} - x_{\text{H}_2\text{dry}} \quad \text{Eq. 1065.655-5}$$

$$x_{\text{dil/exhdry}} = \frac{x_{\text{dil/exh}}}{1 - x_{\text{H}_2\text{Oexh}}} \quad \text{Eq. 1065.655-6}$$

$$x_{\text{int/exhdry}} = \frac{1}{2 \cdot x_{\text{O}_2\text{int}}} \left(\left(\frac{\alpha}{2} - \beta + 2 + 2\gamma \right) (x_{\text{Ccombdry}} - x_{\text{THCdry}}) - (x_{\text{COdry}} - x_{\text{NOdry}} - 2x_{\text{NO}_2\text{dry}} + x_{\text{H}_2\text{dry}}) \right) \quad \text{Eq. 1065.655-7}$$

$$x_{\text{raw/exhdry}} = \frac{1}{2} \left(\left(\frac{\alpha}{2} + \beta + \delta \right) (x_{\text{Ccombdry}} - x_{\text{THCdry}}) + (2x_{\text{THCdry}} + x_{\text{COdry}} - x_{\text{NO}_2\text{dry}} + x_{\text{H}_2\text{dry}}) \right) + x_{\text{int/exhdry}} \quad \text{Eq. 1065.655-8}$$

$$x_{\text{O}_2\text{int}} = \frac{0.209820 - x_{\text{CO}_2\text{intdry}}}{1 + x_{\text{H}_2\text{Ointdry}}} \quad \text{Eq. 1065.655-9}$$

$$x_{\text{CO2int}} = \frac{x_{\text{CO2intdry}}}{1 + x_{\text{H2Ointdry}}} \quad \text{Eq. 1065.655-10}$$

$$x_{\text{H2Ointdry}} = \frac{x_{\text{H2Oint}}}{1 - x_{\text{H2Oint}}} \quad \text{Eq. 1065.655-11}$$

$$x_{\text{CO2dil}} = \frac{x_{\text{CO2dildry}}}{1 + x_{\text{H2Odildry}}} \quad \text{Eq. 1065.655-12}$$

$$x_{\text{H2Odildry}} = \frac{x_{\text{H2Odil}}}{1 - x_{\text{H2Odil}}} \quad \text{Eq. 1065.655-13}$$

$$x_{\text{COdry}} = \frac{x_{\text{COmeas}}}{1 - x_{\text{H2OCOmeas}}} \quad \text{Eq. 1065.655-14}$$

$$x_{\text{CO2dry}} = \frac{x_{\text{CO2meas}}}{1 - x_{\text{H2OCO2meas}}} \quad \text{Eq. 1065.655-15}$$

$$x_{\text{NOdry}} = \frac{x_{\text{NOmeas}}}{1 - x_{\text{H2ONOmeas}}} \quad \text{Eq. 1065.655-16}$$

$$x_{\text{NO2dry}} = \frac{x_{\text{NO2meas}}}{1 - x_{\text{H2ONO2meas}}} \quad \text{Eq. 1065.655-17}$$

$$x_{\text{THCdry}} = \frac{x_{\text{THCmeas}}}{1 - x_{\text{H2OTHCmeas}}} \quad \text{Eq. 1065.655-18}$$

(5) The following example is a solution for $x_{\text{dil/exh}}$, x_{H2Oexh} , and x_{Ccombdry} using the equations in paragraph (c)(4) of this section:

$$x_{\text{dil/exh}} = 1 - \frac{0.184}{1 + \frac{35.38}{1000}} = 0.822 \text{ mol/mol}$$

$$x_{\text{H2Oexh}} = \frac{35.38}{1 + \frac{35.38}{1000}} = 34.18 \text{ mmol/mol}$$

$$x_{\text{Ccombdry}} = 0.025 + \frac{29.3}{1000000} + \frac{47.6}{1000000} - \frac{0.371}{1000} \cdot 0.851 - \frac{0.369}{1000} \cdot 0.172 = 0.0249 \text{ mol/mol}$$

$$x_{\text{H}_2\text{dry}} = \frac{29.3 \cdot (0.034 - 0.012 \cdot 0.851)}{3.5 \cdot \left(\frac{25.2}{1000} - \frac{0.371}{1000} \cdot 0.851 \right)} = 8.5 \mu\text{mol/mol}$$

$$x_{\text{H}_2\text{Oexhdry}} = \frac{1.8}{2} \left(0.0249 - \frac{47.6}{1000000} \right) + 0.018 \cdot 0.851 + 0.017 \cdot 0.172 - \frac{8.5}{1000000} = 0.0353 \text{ mol/mol}$$

$$x_{\text{dil/exhdry}} = \frac{0.822}{1 - 0.034} = 0.851 \text{ mol/mol}$$

$$x_{\text{int/exhdry}} = \frac{1}{2 \cdot 0.206} \left(\left(\frac{1.8}{2} - 0.050 + 2 + 2 \cdot 0.0003 \right) \left(0.0249 - \frac{47.6}{1000000} \right) - \left(\frac{29.3}{1000000} - \frac{50.4}{1000000} - 2 \cdot \frac{12.1}{1000000} + \frac{8.5}{1000000} \right) \right) = 0.172 \text{ mol/mol}$$

$$x_{\text{raw/exhdry}} = \frac{1}{2} \left(\left(\frac{1.8}{2} + 0.050 + 0.0001 \right) \left(0.0249 - \frac{47.6}{1000000} \right) + \left(2 \cdot \frac{47.6}{1000000} + \frac{29.3}{1000000} - \frac{12.1}{1000000} + \frac{8.5}{1000000} \right) \right) + 0.172 = 0.184 \text{ mol/mol}$$

$$x_{\text{O}_2\text{int}} = \frac{0.209820 - 0.000375}{1 + \frac{17.22}{1000}} = 0.206 \text{ mol/mol}$$

$$x_{\text{CO}_2\text{int}} = \frac{0.000375 \cdot 1000}{1 + \frac{17.22}{1000}} = 0.369 \text{ mmol/mol}$$

$$x_{\text{H}_2\text{Ointdry}} = \frac{16.93}{1 - \frac{16.93}{1000}} = 17.22 \text{ mmol/mol}$$

$$x_{\text{CO}_2\text{dil}} = \frac{0.375}{1 + \frac{12.01}{1000}} = 0.371 \text{ mmol/mol}$$

$$x_{\text{H}_2\text{Odildry}} = \frac{11.87}{1 - \frac{11.87}{1000}} = 12.01 \text{ mmol/mol}$$

$$x_{\text{COdry}} = \frac{29.0}{1 - \frac{8.601}{1000}} = 29.3 \text{ mmol/mol}$$

$$x_{\text{CO2dry}} = \frac{24.98}{1 - \frac{8.601}{1000}} = 25.2 \text{ mmol/mol}$$

$$x_{\text{NOdry}} = \frac{50.0}{1 - \frac{8.601}{1000}} = 50.4 \text{ mmol/mol}$$

$$x_{\text{NO2dry}} = \frac{12.0}{1 - \frac{8.601}{1000}} = 12.1 \text{ mmol/mol}$$

$$x_{\text{THCdry}} = \frac{46}{1 - \frac{34.18}{1000}} = 47.6 \text{ mmol/mol}$$

$\alpha = 1.8$
 $\beta = 0.05$
 $\gamma = 0.0003$
 $\delta = 0.0001$

(d) *Carbon mass fraction.* Determine carbon mass fraction of fuel, w_c , using one of the following methods:

(1) You may calculate w_c as described in this paragraph (d)(1) based on measured fuel properties. To do so, you must determine values for α and β in all cases, but you may set γ and δ to zero if the default value listed in Table 1 of this section is zero. Calculate w_c using the following equation:

$$w_c = \frac{1 \cdot M_C}{1 \cdot M_C + \alpha \cdot M_H + \beta \cdot M_O + \gamma \cdot M_S + \delta \cdot M_N} \quad \text{Eq. 1065.655-19}$$

Where:

w_c = carbon mass fraction of fuel.
 M_C = molar mass of carbon.
 α = atomic hydrogen-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.
 M_H = molar mass of hydrogen.
 β = atomic oxygen-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.
 M_O = molar mass of oxygen.
 γ = atomic sulfur-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.
 M_S = molar mass of sulfur.

δ = atomic nitrogen-to-carbon ratio of the mixture of fuel(s) being combusted, weighted by molar consumption.

M_N = molar mass of nitrogen.

Example:

$\alpha = 1.8$
 $\beta = 0.05$
 $\gamma = 0.0003$
 $\delta = 0.0001$
 $M_C = 12.0107$
 $M_H = 1.01$
 $M_O = 15.9994$
 $M_S = 32.065$
 $M_N = 14.0067$

$$w_c = \frac{1 \cdot 12.0107}{1 \cdot 12.0107 + 1.8 \cdot 1.01 + 0.05 \cdot 15.9994 + 0.0003 \cdot 32.065 + 0.0001 \cdot 14.0067}$$

§ 1065.655

40 CFR Ch. I (7-1-11 Edition)

$w_c = 0.8205$

(2) You may use the default values in the following table to determine w_c for a given fuel:

TABLE 1 OF § 1065.655—DEFAULT VALUES OF α , β , γ , δ , AND w_c , FOR VARIOUS FUELS

Fuel	Atomic hydrogen, oxygen, sulfur, and nitrogen-to-carbon ratios $CH\alpha O\beta S\gamma N\delta$	Carbon mass fraction, w_c g/g.
Gasoline	$CH_{1.85}O_0S_0N_0$	0.866
#2 Diesel	$CH_{1.80}O_0S_0N_0$	0.869
#1 Diesel	$CH_{1.93}O_0S_0N_0$	0.861
Liquefied Petroleum Gas	$CH_{2.64}O_0S_0N_0$	0.819
Natural gas	$CH_{3.78}O_{0.016}S_0N_0$	0.747
Ethanol	$CH_3O_{0.5}S_0N_0$	0.521
Methanol	$CH_4O_1S_0N_0$	0.375
Residual fuel blends	Must be determined by measured fuel properties as described in paragraph (d)(1) of this section.	

(e) *Calculated raw exhaust molar flow rate from measured intake air molar flow rate or fuel mass flow rate.* You may calculate the raw exhaust molar flow rate from which you sampled emissions, \dot{n}_{exh} , based on the measured intake air molar flow rate, \dot{n}_{int} , or the measured fuel mass flow rate, \dot{m}_{fuel} , and the values calculated using the chemical balance in paragraph (c) of this section. Note that the chemical balance must be based on raw exhaust gas concentrations. Solve for the chemical balance in paragraph (c) of this section at the same frequency that you update and record \dot{n}_{int} or \dot{m}_{fuel} .

(1) *Crankcase flow rate.* If engines are not subject to crankcase controls under the standard-setting part, you

may calculate raw exhaust flow based on \dot{n}_{int} or \dot{m}_{fuel} using one of the following:

(i) You may measure flow rate through the crankcase vent and subtract it from the calculated exhaust flow.

(ii) You may estimate flow rate through the crankcase vent by engineering analysis as long as the uncertainty in your calculation does not adversely affect your ability to show that your engines comply with applicable emission standards.

(iii) You may assume your crankcase vent flow rate is zero.

(2) *Intake air molar flow rate calculation.* Based on \dot{n}_{int} , calculate \dot{n}_{exh} as follows:

$$\dot{n}_{exh} = \frac{\dot{n}_{int}}{\left(1 + \frac{(x_{int/exhdry} - x_{raw/exhdry})}{(1 + x_{H2Oexhdry})}\right)} \quad \text{Eq. 1065.655-20}$$

Where:

\dot{n}_{exh} = raw exhaust molar flow rate from which you measured emissions.

\dot{n}_{int} = intake air molar flow rate including humidity in intake air.

$\dot{n}_{int} = 3.780 \text{ mol/s}$

$x_{int/exhdry} = 0.69021 \text{ mol/mol}$

$x_{raw/exhdry} = 1.10764 \text{ mol/mol}$

$x_{H2Oexhdry} = 107.64 \text{ mmol/mol} = 0.10764 \text{ mol/mol}$

Example:

$$\dot{n}_{\text{exh}} = \frac{3.780}{\left(1 + \frac{(0.69021 - 1.10764)}{(1 + 0.10764)}\right)}$$

$\dot{n}_{\text{exh}} = 6.066 \text{ mol/s}$

(3) *Fuel mass flow rate calculation.*
Based on \dot{m}_{fuel} , calculate \dot{n}_{exh} as follows:

$$\dot{n}_{\text{exh}} = \frac{\dot{m}_{\text{fuel}} \cdot w_c \cdot (1 + x_{\text{H}_2\text{Oexhdry}})}{M_c \cdot x_{\text{Ccombdry}}} \quad \text{Eq. 1065.655-21}$$

Where:

\dot{n}_{exh} = raw exhaust molar flow rate from which you measured emissions.

\dot{m}_{fuel} = fuel flow rate including humidity in intake air.

Example:

$\dot{m}_{\text{fuel}} = 7.559 \text{ g/s}$

$w_c = 0.869 \text{ g/g}$

$M_c = 12.0107 \text{ g/mol}$

$x_{\text{Ccombdry}} = 99.87 \text{ mmol/mol} = 0.09987 \text{ mol/mol}$

$x_{\text{H}_2\text{Oexhdry}} = 107.64 \text{ mmol/mol} = 0.10764 \text{ mol/mol}$

$$\dot{n}_{\text{exh}} = \frac{7.559 \cdot 0.869 \cdot (1 + 0.10764)}{12.0107 \cdot 0.09987}$$

$\dot{n}_{\text{exh}} = 6.066 \text{ mol/s}$

[73 FR 37331, June 30, 2008, as amended at 73 FR 59334, Oct. 8, 2008; 75 FR 23051, Apr. 30, 2010]

§ 1065.659 Removed water correction.

(a) If you remove water upstream of a concentration measurement, x , or upstream of a flow measurement, n , correct for the removed water. Perform this correction based on the amount of water at the concentration measurement, $x_{\text{H}_2\text{O[emission]meas}}$, and at the flow meter, $x_{\text{H}_2\text{Oexh}}$, whose flow is used to determine the concentration's total mass over a test interval.

(b) When using continuous analyzers downstream of a sample dryer for transient and ramped-modal testing, you must correct for removed water using signals from other continuous analyzers. When using batch analyzers downstream of a sample dryer, you must correct for removed water by using signals either from other batch analyzers or from the flow-weighted average concentrations from contin-

uous analyzers. Downstream of where you removed water, you may determine the amount of water remaining by any of the following:

(1) Measure the dewpoint and absolute pressure downstream of the water removal location and calculate the amount of water remaining as described in §1065.645.

(2) When saturated water vapor conditions exist at a given location, you may use the measured temperature at that location as the dewpoint for the downstream flow. If we ask, you must demonstrate how you know that saturated water vapor conditions exist. Use good engineering judgment to measure the temperature at the appropriate location to accurately reflect the dewpoint of the flow. Note that if you use this option and the water correction in paragraph (d) of this section results in a corrected value that is greater than the measured value, your saturation assumption is invalid and you must determine the water content according to paragraph (b)(1) of this section.